

Supporting Information

Fixation of carbon dioxide into dimethyl carbonate over titanium-based zeolitic thiophene-benzimidazolate framework

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Synthesis of 2-(thiophen-2-yl)-1-((thiophen-2-yl)methyl)-1*H*-benzo[*d*]imidazole

Synthesis of Ti-ZTBF catalyst

Recycling of Ti-ZTBF catalyst

Tentative mechanism for dimethyl carbonate formation from methanol and CO₂

XRD spectra of recycled Ti-ZTBF catalyst

¹H and ¹³C NMR of the product

Synthesis of 2-(thophen-2-yl)-1-((thiophen-2-yl)methyl)-1H-benzo[d]imidazole

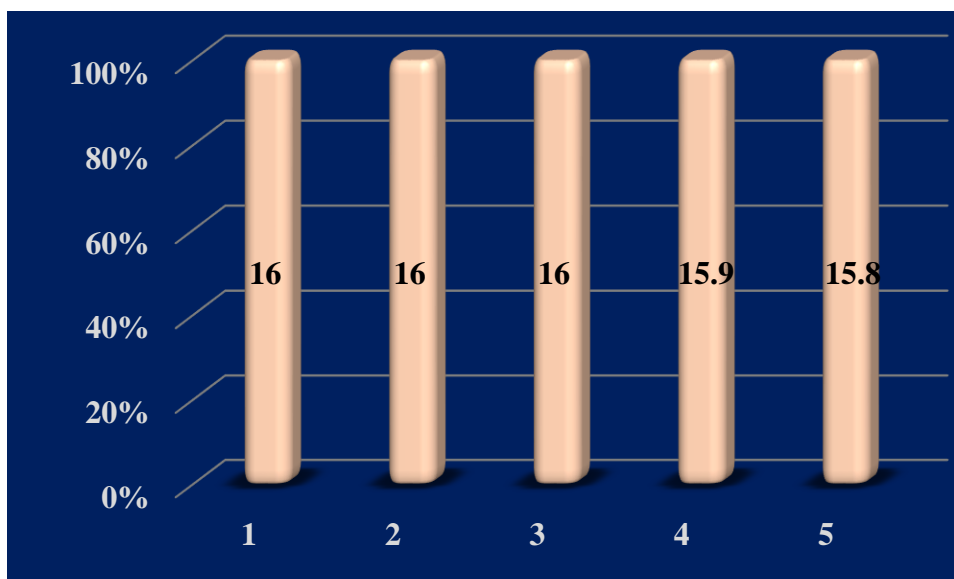
To a mixture of *o*-phenylenediamine (1 mmol; 108 mg) and thiophen-2-carboxyladehhyde (1 mmol; 112 mg) in 5 mL of ethanol, 20 mol % of ammonium chloride was added and the solution was stirred at 80 °C. The progress of reaction was monitored by TLC. After completion of the reaction (6.0 hours), the product was extracted using ethyl acetate, dried over sodium sulfate and concentered under reduced pressure. The product was purified using column chromatography (ethyl acetate: hexane, 1:1) and characterized using GC-MS.

Synthesis of Ti- ZTBF catalyst

Titanium (IV) isobutoxide (2.0 mmol; 680.5 mg), 2-(thophen-2-yl)-1-((thiophen-2-yl)methyl)-1H-benzo[d]imidazole (1 mmol; 296 mg) and dimethylformamide (50 ml) were charged in a pressure reactor. The mixture was heated in 140 °C in an oven for 24 hours. After 24 hours, the reaction was allowed to come down to room temperature. An off-white solid appears in the reaction mixture which was isolated using centrifugation.

3. Recycling of Ti-ZTBF catalyst:

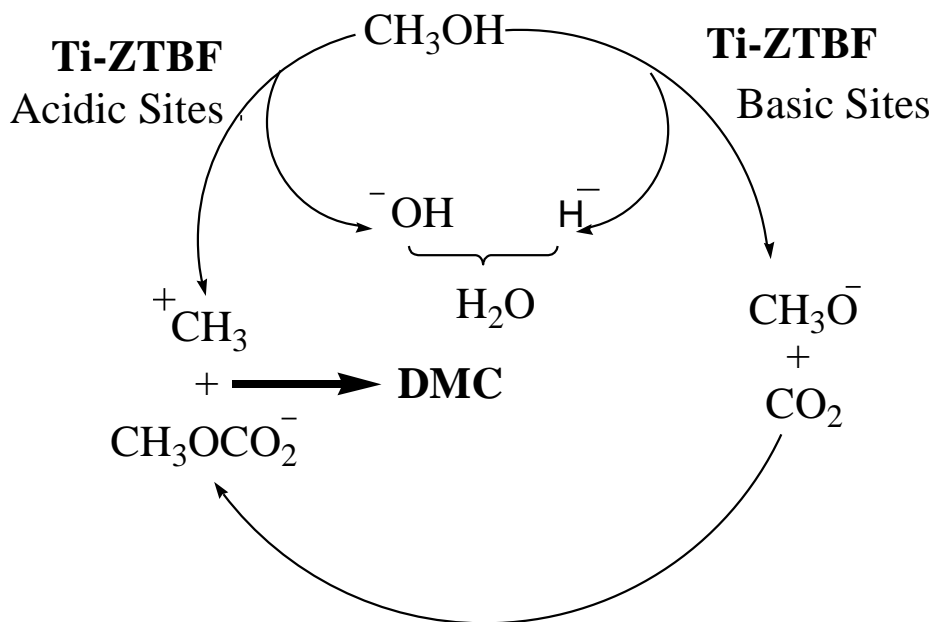
After the completion of each reaction, the Ti-ZTBF catalyst was recovered using a centrifuge, washed with water, dried under vacuum and used for a fresh set of reactants. It was observed that the catalyst remains active even after fifth cycle and could be reused several times without losing activity.



S1. Recycling of Ti-ZTBF catalyst

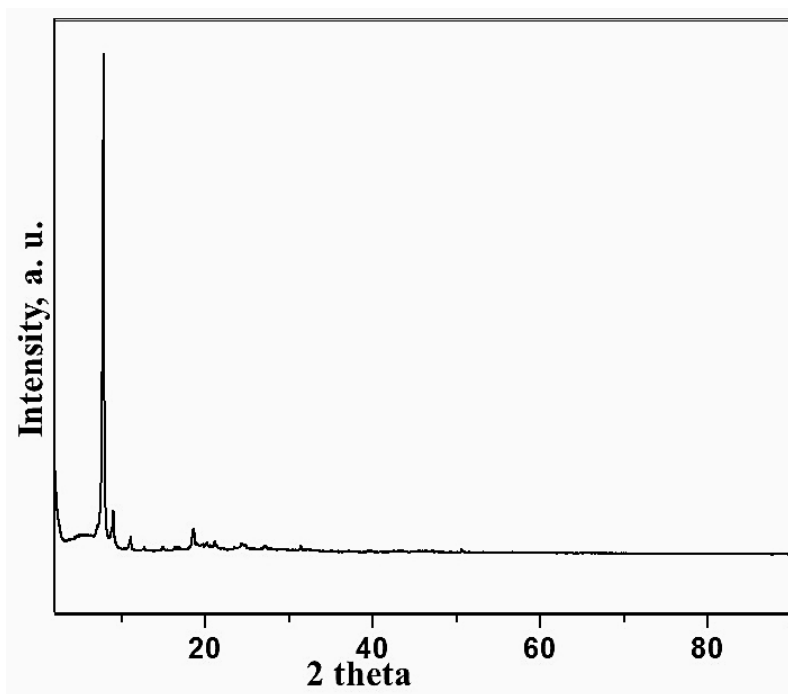
Tentative mechanism for dimethyl carbonate formation from methanol and CO₂

A tentative mechanism is proposed for the synthesis of DMC formation from methanol and CO₂ (S2). Ti-ZTBF catalyst has both, acidic and basic sites which are important factors for the acceleration of the reaction as shown below.



S2. Tentative mechanism for DMC formation from methanol and CO₂

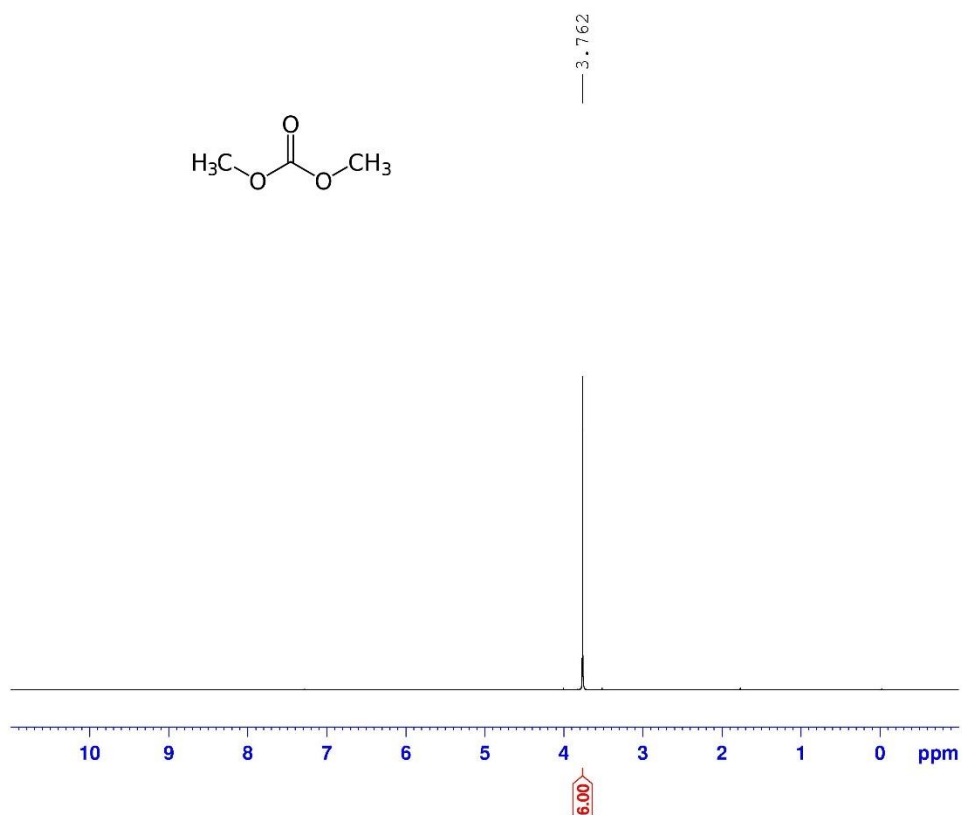
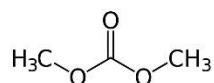
3. XRD spectra of recycled Ti-ZTBF catalyst



S3. XRD spectra of recycled Ti-ZTBF catalyst

¹H and ¹³C NMR of the product

SS-32



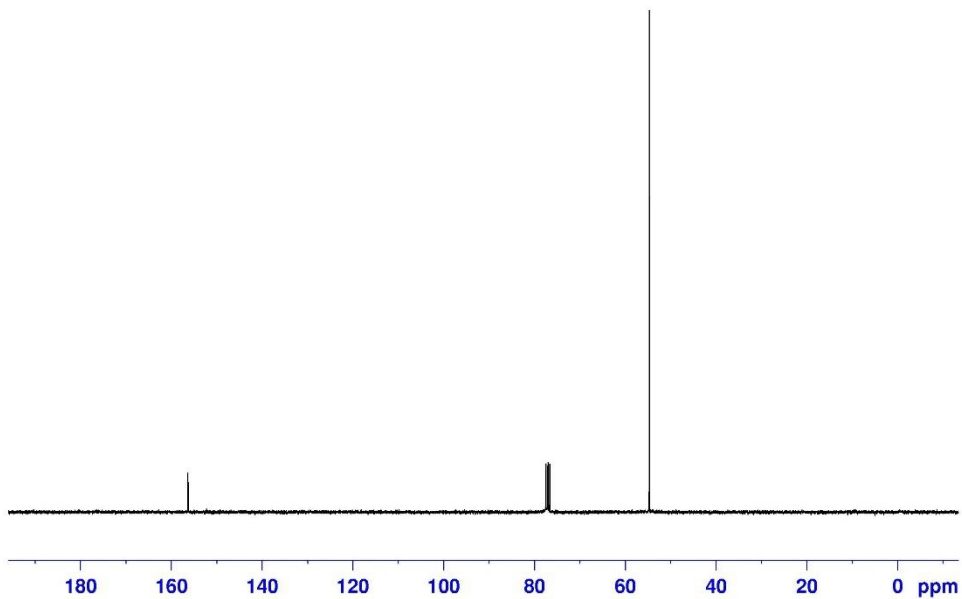
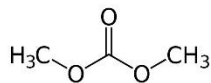
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PROCNO        1
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PULPROG       zg
TD            32768
SOLVENT       CDCl3
NS            16
DS            2
SWH           6188.119 Hz
FIDRES        0.188846 Hz
AQ            2.6477852 sec
RG            181
DW            80.800 usec
DE            6.50 usec
TE            303.0 K
D1            1.0000000 sec
TD0           1
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P1            13.00 usec
PL1           3.20 dB
PL1W          12.02264404 W
SFO1          300.1318534 MHz
SI            32768
SF            300.1300000 MHz
WDW           EM
SSB           0
LB            0.30 Hz
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PC            1.00
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SS-32

156.32

54.69



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TD            32768
SOLVENT       CDCl3
NS            1024
DS            4
SWE           17985.611 Hz
FIDRES        0.548877 Hz
AQ            0.9110282 sec
RG            8192
DW            27.800 usec
DE            6.50 usec
TE            303.0 K
D1            2.0000000 sec
D11           0.0300000 sec
TDC           1

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P1            10.00 usec
PL1           1.80 dB
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SFO1          75.4752953 MHz

----- CHANNEL f2 -----
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PCPD2         80.00 usec
PL2           3.20 dB
PL12          18.98 dB
PL12W         12.02264404 W
PL12W         0.31768745 W
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SI            32768
SF            75.4677490 MHz
WDW           EM
SSB           0
LB            1.00 Hz
GB            0
PC            1.40
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